

### Video « Redox – titration »

Time	Text
00 :09	Now, I will perform the titration of the ascorbic acid by an iodine titrated solution. As ascorbic acid is a reducing agent which will reduce iodine into iodide, this titration is called a redox titration.
00 :27	First, I start the magnetic stirring of the solution in order to get a homogeneous mixture and therefore a homogeneous oxido-reduction.
00 :38	Then, the iodine titrated solution is added drop by drop.
00 :54	As you can see, the iodine is coloured and becomes colourless when it is reduced into iodide by the ascorbic acid.
01 :05	So, the titration is performing by a drop by drop adding of the iodine titrated solution. Previously, I have calculated that the theoretical endpoint volume should be around 5 ml of iodine titrated solution.
01 :22	The oxido-reduction reaction still keeps going. The solution is still colourless.
01 :38	It may be interesting to note that the ascorbic acid titration by an iodine titrated solution is the pharmacopoeia titration.
01 :49	As you can see, I am getting to the end of the titration. All the ascorbic acid has been consumed and you can observe that the iodine is no longer reduced so it stays in solution as coloured diluted iodine.